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LPUNEST Chemistry (BTech) Syllabus PDF

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Chemistry

mass, percentage composition, empirical and molecular formulae; Chemical equations and stoichiometry.

- Real gases, deviation from Ideal behaviour, compressibility factor, van der Waals equation, liquefaction of gases, critical constants. Liquid State: Properties of liquids - vapour pressure, viscosity and surface tension and effect of temperature on them (qualitative treatment only).
- Electrical, magnetic and dielectric properties.
- **UNIT 3:Atomic Structure**

functions; Variation of Ψ and Ψ 2 with r for 1s and 2s orbitals; various quantum numbers (principal, angular momentum and magnetic quantum numbers) and their significance; shapes of s, p and d - orbitals, electron spin and spin quantum number; Rules for filling electrons in orbitals - Aufbau principle, Pauli's exclusion principle and Hund's rule, electronic configuration of elements, extra stability of half-filled and completely filled orbitals.

electron and radii of the different orbits, limitations of Bohr's model; Dual nature of matter, de-Broglie's relationship, Heisenberg uncertainty principle.

UNIT 4: Chemical Bonding and Molecular Structure

> Kossel - Lewis approach to chemical bond formation, concept of ionic and covalent bonds. > Ionic Bonding: Formation of ionic bonds, factors affecting the formation of ionic bonds; calculation of lattice enthalpy.

> Covalent Bonding: Concept of electronegativity, Fajan's rule, dipole moment; Valence Shell Electron Pair Repulsion (VSEPR) theory and shapes of simple

- > Molecular Orbital Theory Its important features, LCAOs, types of molecular orbitals (bonding, antibonding), sigma and pi-bonds, molecular orbital
- **UNIT 5: Chemical Thermodynamics**
- > Fundamentals of thermodynamics: System and surroundings, extensive and intensive properties, state functions, types of processes. > First law of thermodynamics: Concept of work, heat internal energy and enthalpy, heat capacity, molar heat capacity; Hess's law of constant heat

molecular mass using colligative properties; Abnormal value of molar mass, van't Hoff factor and its significance.

UNIT 7: Equilibrium

- > Meaning of equilibrium, concept of dynamic equilibrium.
- > Ionic equilibrium: Weak and strong electrolytes, ionization of electrolytes, various concepts of acids and bases (Arrhenius, Bronsted Lowry and Lewis) and their ionization, acid - base equilibria (including multistage ionization) and ionization constants, ionization of water, pH scale, common ion effect, hydrolysis of salts and pH of their solutions, solubility of sparingly soluble salts and solubility products, buffer solutions.

> Equilibria involving physical processes: Solid -liquid, liquid - gas and solid - gas equilibria, Henry's law, general characteristics of equilibrium involving physical

cell reactions, emf of a Galvanic cell and its measurement; Nernst equation and its applications; Relationship between cell potential and Gibbs' energy change; Dry cell and lead accumulator; Fuel cells; Corrosion and its prevention. **UNIT 9: Chemical Kinetics**

> Rate of a chemical reaction, factors affecting the rate of reactions: concentration, temperature, pressure and catalyst; elementary and complex reactions,

order and molecularity of reactions, rate law, rate constant and its units, differential and integral forms of zero and first order reactions, their characteristics

and half - lives, effect of temperature on rate of reactions - Arrhenius theory, activation energy and its calculation, collision theory of bimolecular gaseous

isotherms, adsorption from solutions.

reactions (no derivation).

and associated colloids (micelles), preparation and properties of colloids - Tyndall effect, Brownian movement, electrophoresis, dialysis, coagulation and flocculation; Emulsions and their characteristics.

> Modem periodic law and present form of the periodic table, s, p, d and f block elements, periodic trends in properties of elements, atomic and ionic radii,

> Colloidal state - distinction among true solutions, colloids and suspensions, classification of colloids - lyophilic, lyophobic; multi molecular, macromolecular

> Catalysis - Homogeneous and heterogeneous, activity and selectivity of solid catalysts, enzyme catalysis and its mechanism.

UNIT 12: General Principles and Process of Isolation of Metals > Modes of occurrence of elements in nature, minerals, ores; Steps involved in the extraction of metals - concentration, reduction (chemical. and electrolytic

> Position of Hydrogen in periodic table, isotopes, preparation, properties and uses of Hydrogen; Physical and chemical properties of water and heavy water;

Structure, preparation, reactions and uses of Hydrogen peroxide; Classification of Hydrides - ionic, covalent and interstitial; Hydrogen as a fuel.

methods) and refining with special reference to the extraction of Al, Cu, Zn and Fe; Thermodynamic and electrochemical principles involved in the extraction

> Group 1 and Group 2 Elements

UNIT 14: s - Block Elements (Alkali and Alkaline Earth Metals)

UNIT 11: Classification of Elements and Periodicity in Properties

ionization enthalpy, electron gain enthalpy, valence, oxidation states and chemical reactivity.

> Preparation and properties of some important compounds - Sodium carbonate, Sodium chloride, Sodium hydroxide and Sodium hydrogen carbonate; Industrial uses of Lime, Limestone, Plaster of Paris and cement; Biological significance of Na, K, Mg and Ca.

> General introduction, electronic configuration and general trends in physical and chemical properties of elements, anomalous properties of the first element

> Group wise study of the p - block elements > Group - 13

chloride and alums.

> Group - 14

> Group - 16

> Group - 18

> Group 13 to Group 15 Elements

of each group, diagonal relationships.

unique behaviour of the first element in each group.

of metals.

> Tendency for catenation; Structure, properties and uses of allotropes and oxides of Carbon, Silicon tetrachloride, Silicates, Zeolites and Silicones. > Group - 15 > Properties and uses of Nitrogen and Phosphorus; Allotrophic forms of Phosphorus; Preparation, properties, structure and uses of Ammonia, Nitric acid,

Phosphine and Phosphorus halides, (PCI3, PCI5); Structures of oxides and oxoacids of Nitrogen and Phosphorus.

> Preparation, properties and uses of Boron and Aluminium; Structure, properties and uses of Borax, Boric acid, Diborane, Boron tri-fluoride, Aluminium

dioxide, Sulphuric acid (including its industrial preparation); Structures of oxoacids of Sulphur. > Group - 17

> Group wise study of the p - block elements

unique behaviour of the first element in each group.

> Actinoids - Electronic configuration and oxidation states.

UNIT 18: Co-ordination Compounds

Occurrence and uses of noble gases; Structures of fluorides and oxides of xenon.

UNIT 20: Purification and Characterization of Organic Compounds

> Qualitative analysis - Detection of nitrogen, Sulphur, phosphorus and halogens.

UNIT 16: p - Block Elements

alloy formation; Preparation, properties and uses of K2Cr2O7 and KMnO4. > Inner Transition Elements > Lanthanoids - Electronic configuration, oxidation states, chemical reactivity and lanthanoid contraction.

> General introduction, electronic configuration, occurrence and characteristics, general trends in properties of the first row transition elements - physical

properties, ionization enthalpy, oxidation states, atomic radii, colour, catalytic behaviour, magnetic properties, complex formation, interstitial compounds,

- **UNIT 19: Environmental Chemistry** > Environmental pollution - Atmospheric, water and soil. > Atmospheric pollution - Tropospheric and stratospheric
- > Nomenclature (Trivial and IUPAC) > Covalent bond fission - Homolytic and heterolytic: free radicals, carbocations and carbanions; stability of carbocations and free radicals, electrophiles and nucleophiles.

> Electronic displacement in a covalent bond - Inductive effect, electromeric effect, resonance and hyper conjugation.

C - and those containing Halogens, Oxygen, Nitrogen and Sulphur; Homologous series; Isomerism - structural and stereoisomerism.

> Purification - Crystallization, sublimation, distillation, differential extraction and chromatography - principles and their applications.

> Quantitative analysis (basic principles only) - Estimation of Carbon, Hydrogen, Nitrogen, Halogens, Sulphur, Phosphorus.

> Calculations of empirical formulae and molecular formulae; Numerical problems in organic quantitative analysis.

Alcohols, Phenols and Ethers Alcohols: Identification of primary, secondary and tertiary Alcohols; mechanism of dehydration. > Phenols: Acidic nature, electrophilic substitution reactions: halogenation, nitration and sulphonation, Reimer - Tiemann reaction.

> General introduction and classification of polymers, general methods of polymerization - addition and condensation, copolymerization;

> Natural and synthetic rubber and vulcanization; some important polymers with emphasis on their monomers and uses - Polythene, Nylon, Polyester and

> Chemistry involved in the preparation of the following: Inorganic compounds: Mohr's salt, potash alum. Organic compounds: Acetanilide, pnitroacetanilide,

- > Amines: Nomenclature, classification, structure, basic character and identification of primary, secondary and tertiary amines and their basic character. > Diazonium Salts: Importance in synthetic organic chemistry. **UNIT 26: Polymers**
- Enthalpy of neutralization of strong acid and strong base.
- meaning and common examples. Chemicals in food - Preservatives, artificial sweetening agents - common examples.
- > Chemistry involved in the titrimetric excercises Acids bases and the use of indicators, oxalic-acid vs KMnO4, Mohr's salt vs KMnO4. > Chemical principles involved in the qualitative salt analysis: Cations - Pb2+, Cu2+, Al3+, Fe3+, Zn2+, Ni2+, Ca2+, Ba2+, Mg2+, NH4+. Anions- CO3 2-, S2-, SO4 2-, NO2-, NO3-, CI -, Br-, I-. (Insoluble salts excluded).
- > Enthalpy of solution of CuSO4
- UNIT 28: Chemistry in Everyday Life
- > Chemicals in medicines Analgesics, tranquilizers, antiseptics, disinfectants, antimicrobials, antifertility drugs, antibiotics, antacids, antihistamins their
- > General introduction and importance of biomolecules. > Carbohydrates - Classification: aldoses and ketoses; monosaccharides (glucose and fructose), constituent monosaccharides of oligosacchorides (sucrose, lactose, maltose) and polysaccharides (starch, cellulose, glycogen).
 - UNIT 30: Stratospheric pollution Formation and breakdown of ozone, depletion of ozone layer its mechanism and

> Proteins - Elementary Idea of amino acids, peptide bond, polypeptides; Proteins: primary, secondary, tertiary and quaternary structure (qualitative idea only),

- > Solid State: Classification of solids: molecular, ionic, covalent and metallic solids, amorphous and crystalline solids (elementary idea); Bragg's Law and its applications; Unit cell and lattices, packing in solids (fcc, bcc and hcp lattices), voids, calculations involving unit cell parameters, imperfection in solids;
- > Discovery of sub-atomic particles (electron, proton and neutron); Thomson and Rutherford atomic models and their limitations; Nature of electromagnetic radiation, photoelectric effect; Spectrum of hydrogen atom, Bohr model of hydrogen atom - its postulates, derivation of the relations for energy of the
- > Elementary ideas of quantum mechanics, quantum mechanical model of atom, its important features, concept of atomic orbitals as one electron wave
- molecules.
- > Quantum mechanical approach to covalent bonding: Valence bond theory Its important features, concept of hybridization involving s, p and d orbitals; Resonance.
- electronic configurations of homonuclear diatomic molecules, concept of bond order, bond length and bond energy. > Elementary idea of metallic bonding. Hydrogen bonding and its applications.
- summation; Enthalpies of bond dissociation, combustion, formation, atomization, sublimation, phase transition, hydration, ionization and solution.

> Different methods for expressing concentration of solution - molality, mole fraction, percentage (by volume and mass both), vapour pressure of

dilute solutions - relative lowering of vapour pressure, depression of freezing point, elevation of boiling point and osmotic pressure; Determination of

solutions and Raoult's Law - Ideal and non-ideal solutions, vapour pressure - composition, plots for ideal and non-ideal solutions; Colligative properties of

- > Second law of thermodynamics: Spontaneity of processes; ΔS of the universe and ΔG of the system as criteria for spontaneity, ΔGo (Standard Gibbs energy change) and equilibrium constant.
- **UNIT 6: Solutions**

UNIT 8: Redox Reactions and Electrochemistry

processes. > Equilibria involving chemical processes: Law of chemical equilibrium, equilibrium constants (Kp and Kc) and their significance, significance of ΔG and ΔGo in

chemical equilibria, factors affecting equilibrium concentration, pressure, temperature, effect of catalyst; Le Chatelier's principle.

Kohlrausch's law and its applications. > Electrochemical cells - Electrolytic and Galvanic cells, different types of electrodes, electrode potentials including standard electrode potential, half - cell and

> Electronic concepts of oxidation and reduction, redox reactions, oxidation number, rules for assigning oxidation number, balancing of redox reactions.

> Eectrolytic and metallic conduction, conductance in electrolytic solutions, specific and molar conductivities and their variation with concentration:

- **UNIT 10: Surface Chemistry** > Adsorption - Physisorption and chemisorption and their characteristics, factors affecting adsorption of gases on solids - Freundlich and Langmuir adsorption
- UNIT 13: Hydrogen
- UNIT 15: p Block Elements

> General Introduction: Electronic configuration and general trends in physical and chemical properties of elements across the periods and down the groups;

> Group 16 to Group 18 Elements: > General Introduction: Electronic configuration and general trends in physical ad chenmical properties of elements across the periods and down the groups;

> Preparation, properties, structures and uses of dioxygen and ozone; Allotropic forms of Sulphur; Preparation, properties, structures and uses of Sulphur

> Preparation, properties and uses of hydrochloric acid; Trends in the acidic nature of hydrogen halides; Structures of Interhalogen compounds and oxides and

UNIT 17: d - and f - Block Elements: > Transition Elements

oxoacids of halogens.

> Introduction to co-ordination compounds, Werner's theory; ligands, co-ordination number, denticity, chelation; IUPAC nomenclature of mononuclear coordination compounds, isomerism; Bonding-Valence bond approach and basic ideas of Crystal field theory, colour and magnetic properties; Importance of co-ordination compounds (in qualitative analysis, extraction of metals and in biological systems).

> Tropospheric pollutants - Gaseous pollutants: Oxides of Carbon, Nitrogen and Sulphur, hydrocarbons; their sources, harmful effects and prevention; Green

> Tetravalency of Carbon; Shapes of simple molecules - hybridization (s and p); Classification of organic compounds based on functional groups: - C = C - , - C =

> Aromatic hydrocarbons - Nomenclature, benzene - structure and aromaticity; Mechanism of electrophilic substitution; halogenation, nitration, Friedel -

house effect and Global warming; Acid rain; Particulate pollutants: Smoke, dust, smog, fumes, mist; their sources, harmful effects and prevention.

UNIT 21: Some Basic Principles of Organic Chemistry-I

UNIT 22: Hydrocarbons

> Ethers: Structure.

> Aldehyde and Ketones

Bakelite.

UNIT 27: Practical Chemistry

aniline yellow, iodoform.

Some Basic Principles of Organic Chemistry-II:

Organic Compounds Containing Halogens:

> Uses; Environmental effects of chloroform, iodoform, freons and DDT.

UNIT 23: Organic Compounds Containing Oxygen-I

> General methods of preparation, properties, reactions and uses.

UNIT 24: Organic Compounds Containing Oxygen-II

> General methods of preparation, properties, reactions and uses.

> Carboxylic Acids: Acidic strength and factors affecting it.

UNIT 25: Organic Compounds Containing Nitrogen:

> General methods of preparation, properties, reactions and uses.

> Alkanes - Conformations: Sawhorse and Newman projections (of ethane); Mechanism of halogenation of Alkanes. > Alkenes - Geometrical isomerism; Mechanism of electrophilic addition: addition of hydrogen, halogens, water, hydrogen halides (Markownikoff's and peroxide effect); Ozonolysis, oxidation, and polymerization.

> Alkynes - Acidic character; Addition of hydrogen, halogens, water and hydrogen halides; Polymerization.

Craft's alkylation and acylation, directive influence of functional group in mono-substituted benzene.

> General methods of preparation, properties and reactions; Nature of C-X bond; Mechanisms of substitution reactions.

> Classification, isomerism, IUPAC nomenclature, general methods of preparation, properties and reactions.

> Common types of organic reactions - Substitution, addition, elimination and rearrangement.

- > Nature of carbonyl group; Nucleophilic addition to >C=O group, relative reactivities of aldehydes and ketones; Important reactions such as Nucleophilic addition reactions (addition of HCN, NH3 and its derivatives), Grignard reagent; oxidation; reduction (Wolff Kishner and Clemmensen); acidity of - Hydrogen, aldol condensation, Cannizzaro reaction, Haloform reaction; Chemical tests to distinguish between Aldehydes and Ketones.
- > Cleansing agents Soaps and detergents, cleansing action. **UNIT 29: Bio Molecules**
- > Nucleic Acids Chemical constitution of DNA and RNA. Biological functions of nucleic acids.

denaturation of proteins, enzymes.

> Vitamins - Classification and functions.

effects.

> Soil pollution - Major pollutants such as: Pesticides (insecticides, herbicides and fungicides), their harmful effects and prevention.

> Water Pollution - Major pollutants such as, pathogens, organic wastes and chemical pollutants; their harmful effects and prevention.

- Strategies to control environmental pollution

- Classification of matter into solid, liquid and gaseous states. > Gaseous State: Measurable properties of gases; Gas laws - Boyle's law, Charle's law, Graham's law of diffusion, Avogadro's law, Dalton's law of partial pressure; Concept of Absolute scale of temperature; Ideal gas equation, Kinetic theory of gases; Concept of average, root mean square and most probable velocities;
- **UNIT 2: States of Matter**

- UNIT 1: Some Basic concepts in Chemistry
- Matter and its nature, Dalton's atomic theory; Concept of atom, molecule, element and compound; Physical quantities and their measurements in Chemistry, precision and accuracy, significant figures, S.I. Units, dimensional analysis; Laws of chemical combination; Atomic and molecular masses, mole concept, molar